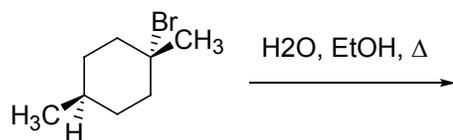


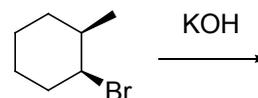
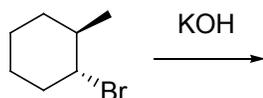
Chem 30A- Week 9

Substitution vs. Elimination

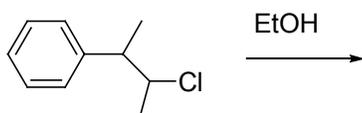
1.



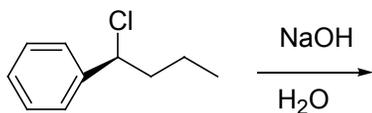
2.



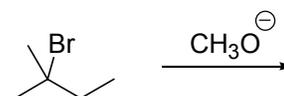
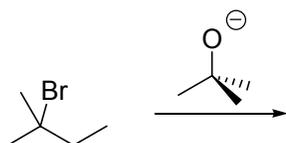
3.



4.

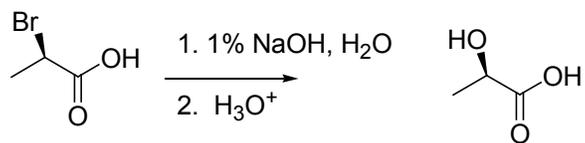


5.

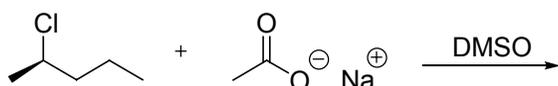


Chapters 8

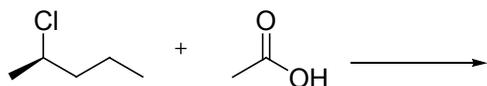
6. Propose a mechanism:



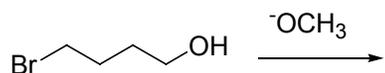
7.



8.



9.

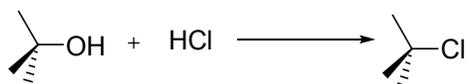


Hint: 2 possible products

10. Explain why t-butanol will not react with potassium chloride, but will react with concentrated HCl to form 2-chloro-2-methylpropane.

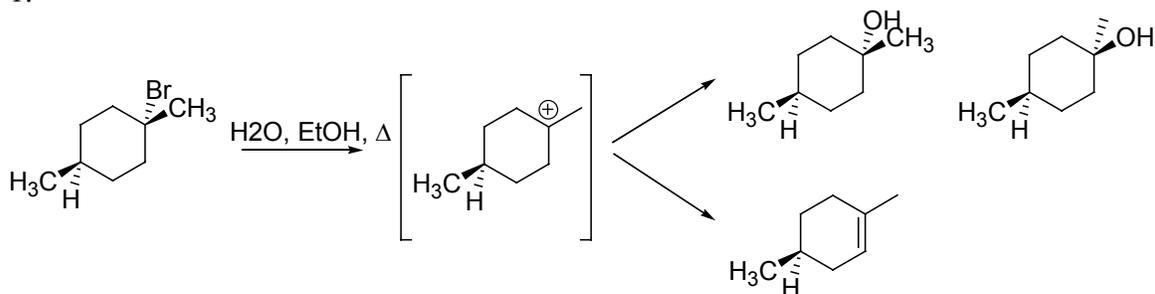


But

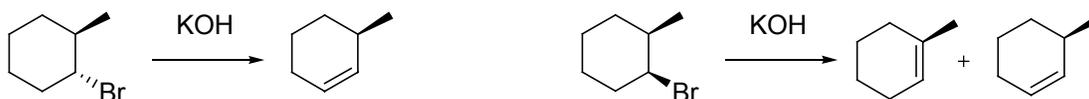


Chem 30A- Week 9 KEYSubstitution vs. Elimination

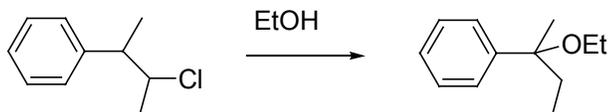
1.



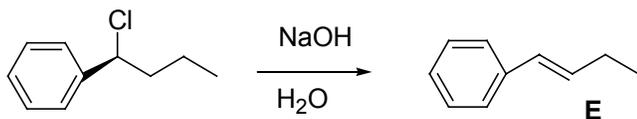
2.



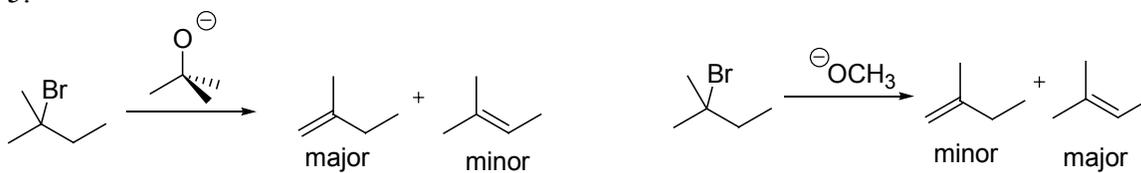
3.



4.

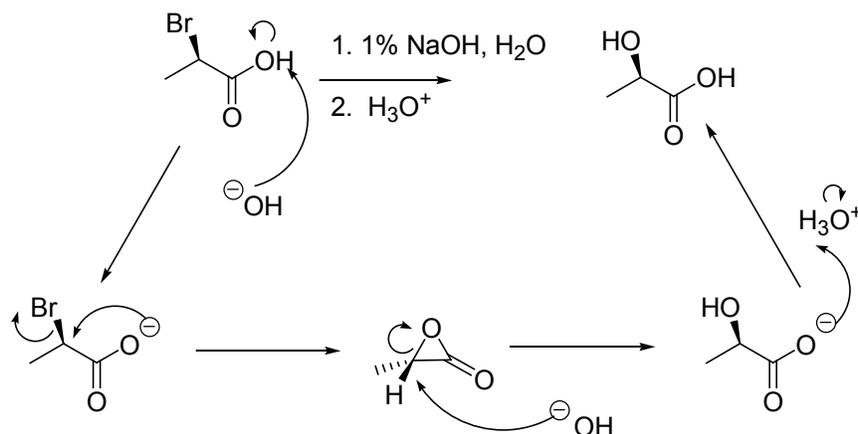


5.

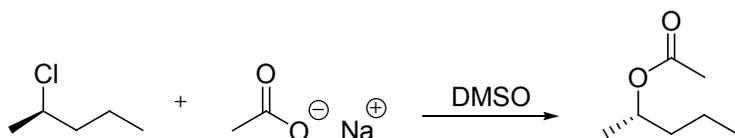


Chapters 8

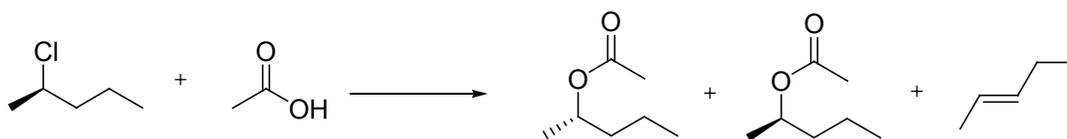
7. Propose a mechanism:



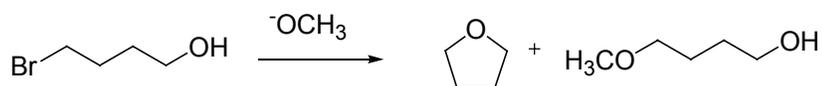
7.



8.



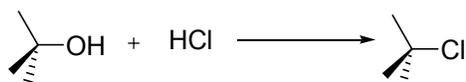
9.



11. Explain why t-butanol will not react with potassium chloride, but will react with concentrated HCl to form 2-chloro-2-methylpropane.



But



OH is not a good leaving group, but HCl can protonate to make $\begin{matrix} \text{H} \\ | \\ \text{---O}^+ \\ | \\ \text{H} \end{matrix}$ which is. Now a S_N1 reaction will proceed.