## Chem 14D – Spring 2011

## pK<sub>a</sub> Table

## (see Vollhardt: p. 60 in 5<sup>th</sup> Edition or 6<sup>th</sup> Edition)

Name of Acid	<b>Chemical Formula</b>	p <i>K</i> <sub>a</sub>
Hydrogen iodide	HI	-10.0
Hydrogen bromide	HBr	-9.0
Hydrogen chloride	HC1	-8.0
Sulfuric acid	$H_2SO_4$	-3.0
Hydronium ion	$H_3O^+$	-1.7
Nitric acid	HNO <sub>3</sub>	-1.4
Methanesulfonic acid	CH <sub>3</sub> SO <sub>3</sub> H	-1.2
Hydrogen fluoride	HF	3.2
Acetic acid	CH <sub>3</sub> COOH	4.7
Hydrogen cyanide	HCN	9.2
Ammonium ion	$\mathrm{NH_4}^+$	9.3
Methanethiol	CH <sub>3</sub> SH	10.0
Methanol	CH <sub>3</sub> OH	15.5
Water	H <sub>2</sub> O	15.7
Ammonia	NH <sub>3</sub>	35
Methane	CH <sub>4</sub>	~50