

CHEM 30A, W05, SET #1

1) Draw Lewis structures for the following compounds:

a) CH_3OH

b) CH_3Br

c) HCN

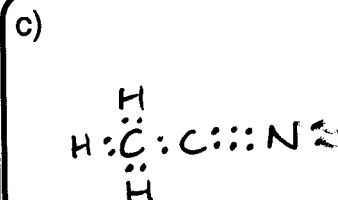
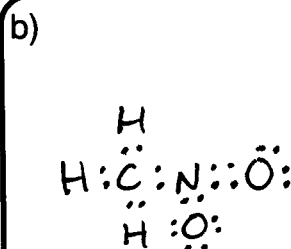
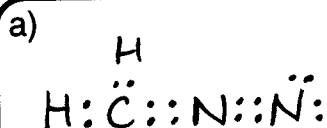
2) Circle the **most** electronegative element in each compound:

a) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$

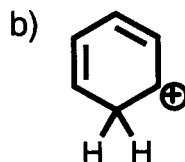
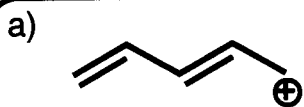
b) CH_3NHCH_3

c) $\text{FCH}_2\text{CH}_2\text{OH}$

3) Calculate the formal charge on any **heteroatom**:



4) Draw resonance structures for the following compounds:



5) Draw the simplified line structures for the following compounds:

a) $(\text{CH}_3\text{CH}_2)_2\text{CHCH}_2\text{CH}_3$

